مراجعة

كيمياء عامة CH3

لطلاب السنة التحضيرية بجامعة الملك خالد بأبها – المحالة

2015

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لا تنسى زيارة المدونة http://kku-a.blogspot.com

و نشر المدونة بين الطلاب لتعم الفائدة

1) Light behaves as :

a) wave b) partical c) both a and b

d) neither a nor b

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2) The wavelength of sodium lamps light is 589nm w is frequency of this radiation ? $C = 3 \times 10^{8} m s^{-1}$

a) $5.09 \times 10^{12}\ sec^{-1}$ b) $5.09 \times 10^{5}\ sec^{-1}$ c) 5.09×10^{14}

d) $3 \times 10^{5}\ sec^{-1}$

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3) Pauli Exclusion principle State that

a) Electrons accupy degenerate orbitals ~~stg~~ singly to The maximum extent

b) NO two electrons in an atom Can have the same set four quantum numbers

c) Electrons travels in waves like light

d) Electron of H atom move in a circular Path

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4) The distance between identical points on success waves is Called :

a) Frequency b) period time c) wavenumber

d) Wave length

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7) The electromagnetic radiation with longest wavelength is

a) Radio wave b) infra-red c) x-rays d) U

5) Which of the following type of radiation has the highest energy :

a) U.V b) Micro wave c) infra-red

d) X-rays

7) The sublevel orbital which can be accupied with maximum 14 electrons is

a) 3d b) 4f c) 5S d) 6P

8) Which set of quantum numbers can be used to characterize the sublevel (3P)

a) $n = 4 \quad l = 1 \quad m_l = 0 \quad m_s = +\frac{1}{2}$

b) $n = 3 \quad l = 2 \quad m_l = 1 \quad m_s = -\frac{1}{2}$

c) $n = 3 \quad l = 1 \quad m_l = 1 \quad m_s = -\frac{1}{2}$

d) $n = 3 \quad l = 0 \quad m_l = 0 \quad m_s = +\frac{1}{2}$

1) the electronic configuration of Cu_{29} is

a) $_{18}Ar\ 4s^2\ 3d^9$ b) $_{18}Ar\ 4s^1\ 3d^9$ c) $_{18}Ar\ 3d^{10}\ 4p^1$

d) $Ar\ 4s^1\ 3d^{10}$

11) which of sublevel is characterized by the set of quant numbers $n = 4$ $l = 3$ $m_l = 2$ $m_s = \frac{1}{2}$

a) 3P b) 4d c) 4P d) 4f

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10) How many unpaired electrons are there in an atom of oxygen ($^{O}_{8}$)

a) 1 b) 2 c) 3 d) 4

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12) what is electronic configuration for Ni^{2+} {Ni=28

a) $_{18}Ar\ 4s^2 3d^8$ b) $_{18}Ar\ 4s^2 3d^6$ c) $_{18}Ar\ 3d^8$

d) $_{18}Ar 4s^2 3d^{10}$

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13) The symbol of the element of lowest atomic numb whose ground state has ~~Five~~ Six 3d electron

a) Fe 26 b) Co 27 c) Cu 29 d) Zn 30

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14) write the electronic configuration of Mg^{++} (Mg =12

a) $1s^2 2s^2 2p^6 3s^2$ b) $1s^2 2s^2 2p^6$ c) $1s^2 2s^2 2p^6 3s$ 3+

d) [Ne] $3s^2$

15) The ..... number determines the Shape of the or

a) l b) n c) m_l d) m

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6) The number of orbitals in d ~~Subshell~~ SUbleVed

a) 10 b) 14 c) 6 d) 5

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7) The ion that is isoelectronic with Ne (10) is

a) Na^+ b) K^+ c) Ca^{++} d) Li^+

Na=11 K=19 Ca=20 Li=3

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18) The energy required to remove an electron from a gas atom or ion

a) electronegativaty b) ionizatt on Energy

c) electron affinity d) none of them

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19) Which of the following atoms has The greatest electronegativaty

a) I b) Br c) Cl d) F

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20) Planck suggested all energy gained or lost b an atom must be some integrel multiple of a minimum amount of energy called

a) electron b) quantum c) nucleus

Which Pair is given in The Correct order of increas Size With respect to ionic

a) $Na^+ > Na$ b) $Cr^{+6} > Cr^{+3}$ c) $Cl > Cl^-$

d) $Ca > Mg$

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22) What is the Valence Shell electrons Configuratio of Phosphorus $_{15}P$?

a) $3s^2 3p^3$ b) $3p^3$ c) $1s^2 2s^2 2p^6$ d) $2s^2$

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23) The energy required to exict electron from ground state ($n=1$) to $n=3$ is

a) 2.73×10^{-18} J b) 4.70×10^{-18} J c) 1.94×10^{-18} J d) $1.7 \times$

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24) The ------ quantum number determines the Orientation of electron (Spin) in the magnetic field

a) n b) m_s c) m_ℓ d) ℓ

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25) Which of the following elements has the largest atomic radius?

a) $_9F$ b) $_5B$ c) $_4Be$ d) $_6C$

1	2	3	4	5	6	7	8	9	10	11	12
c	c	b	d	a	d	b	c	d	d	b	c
13	14	15	16	17	18	19	20	21	22	23	24	25
a	b	a	d	a	b	d	b	d	a	c	b	c

تمنياتي بالتفوق
لاتنسونا من الدعاء
صلاح S.A

ملحوظة: اذا اردت ان تقارن بين طاقتي مجالين
تم حساب (n+l) لكل مجال
الاكبر هو أعلى طاقة
اذا تساوا ننظر الى n فقط n أكبر أعلى طاقة

4s 3d

4s n+l = 4+0 = 4

3d n+l = 3+2 = 5

∴ 3d > 4s

6p 7s

n+l = 7 n+l = 7 7s > 6p